Formula Writing Review #4A
Dr. Slotsky
Write formulas for the following compounds

1) Iron (III) Oxide _____

2) Iron (II) Sulfide _____

3) Copper (I) Iodide _____

4) Copper (I) Oxide _____

5) Copper (II) Oxide _____

6) Lead (II) Chloride _____

7) Lead (IV) Oxide _____

8) Antimony (V) Fluoride _____

9) Mercury (II) Oxide _____

10) Titanium (IV) Chloride _____

EXAMPLE: Tin (IV) Oxide is SnO_2

The symbol for Tin is Sn, and Oxide indicates oxygen, O. Consulting our periodic tables, we can see that Oxygen forms the negative ion O^{-2}. Tin forms both Sn^{+2} and Sn^{+4} ions – the Roman numeral tells us that, here, we have Sn^{+4}. In order for the +4 charge of the Sn ion to be balanced, we need 2 of the -2 Oxide ions, giving a formula of SnO_2.