

Formula Writing Review #5

Dr. Slotsky

Write names for the following compounds

- 1) $(\text{NH}_4)_2\text{CO}_3$ _____
- 2) Na_3PO_4 _____
- 3) K_2SO_4 _____
- 4) K_2SO_3 _____
- 4) $(\text{NH}_4)_2\text{S}$ _____
- 5) LiClO_4 _____
- 6) LiClO_3 _____
- 7) NaHCO_3 _____
- 8) NH_4NO_3 _____
- 9) CaCO_3 _____
- 10) AlPO_4 _____

EXAMPLE: $\text{Ca}(\text{CN})_2$ is Calcium cyanide

We are given the formula $\text{Ca}(\text{CN})_2$. Consulting the table of common polyatomic ions on page 178 of your textbook, we notice that CN^- is a polyatomic ion known as “cyanide” – famous as a poison, perhaps, to some who read murder mysteries. We know that Ca means the element Calcium, which forms the Ca^{+2} ion (ref p. 160). Putting the positive ion first, as is the rule for writing chemical names, we write “Calcium cyanide”.