

Name _____

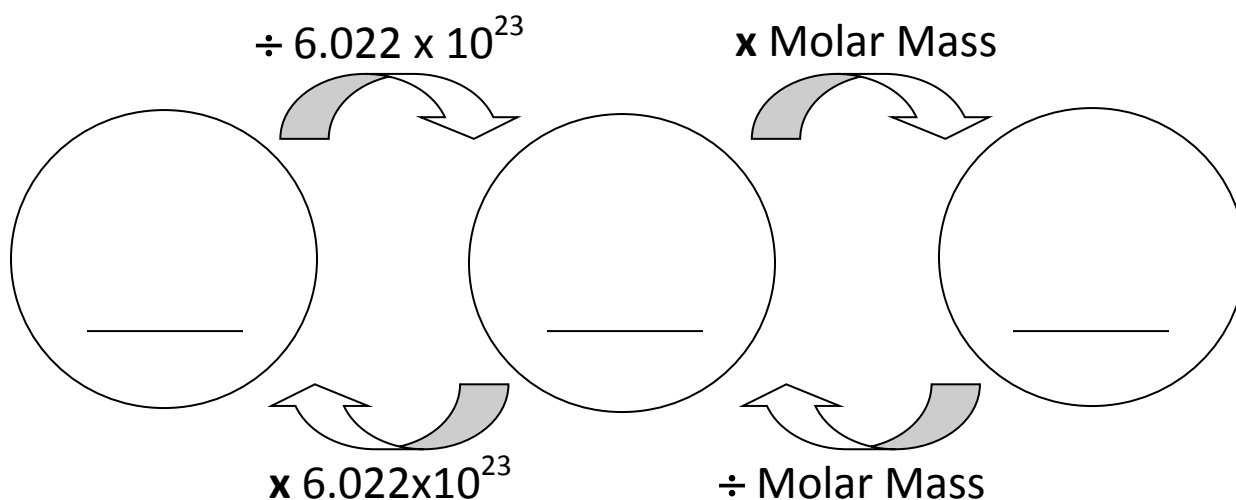
Introduction to The Mole

Part 1: What is a Mole?

A mole is the SI unit for _____ of a pure substance. It is equal to _____ atoms or molecules. A mole is also equal to one _____ mass of a substance. The molar mass of an element is a mass in _____ equal to the atomic _____ for that element in the Periodic Table. The molar mass of a _____ is equal, in grams, to the _____ of the atomic masses of the atoms which make up the compound.

Word Bank: 6.022×10^{23} , molar, grams, amount, mass, compound, element, sum

Part 2: "Moles" have TWO Jobs: Particles \leftrightarrow Moles and Moles \leftrightarrow Grams



Word Bank: Mass in Grams, Amount in Moles, Number of Particles

Students are often confused by moles. The rule is that you use 6.022×10^{23} (Avogadro's Number) when you're calculating the number of _____ or _____, which is only really used in textbook problems. When you're calculating a mass in _____, which is useful for planning _____ experiments, use your _____ instead! Moles provide a _____ between number of atoms and mass in grams. Keep these two things straight and moles will be much easier!

Word Bank: lab, Periodic Table, grams, connection, atoms, molecules

I'm the hardest concept in high school chemistry!

