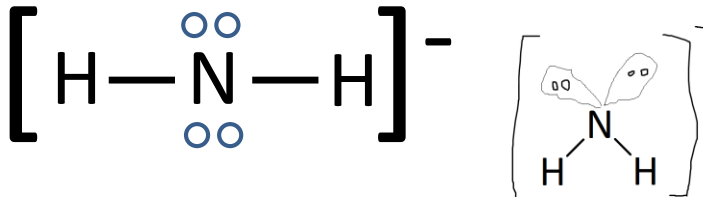


Name \_\_\_\_\_

## Shapes of Molecules – Practice

Draw a Lewis structure of each molecule or ion – if there are multiple resonance structures, show them all. The central atom of each molecule or ion will be underlined> as a hint. Use the Lewis structure (and sketch, if necessary) to classify the geometry of each molecule as one of the following: Linear, Bent, Triangular, Pyramidal, Tetrahedral.

EXAMPLE: NH<sub>2</sub><sup>-</sup>



Geometry: BENT

F<sub>2</sub>O

CS<sub>2</sub>

BH<sub>3</sub> (Note: boron has an incomplete octet in this molecule)

AsH<sub>3</sub>

SiH<sub>4</sub>

(HONORS) NO<sub>3</sub><sup>-</sup> (Note: there are 3 resonance structures. The ‘real’ structure of the ion is an average between them)