Step 1 is to prepare Fe(OH)3 (actually, hydrated Fe2O3). The easiest method, probably, is by acting with NaHCO3 or NaOH on FeCl3 solution. I, however, used a longer way:

- Precipitate FeCO3 by reacting FeSO4 with NaHCO3. It has dirty greenish color.
- Oxidize FeCO3 to Fe(OH)3 by bleach: add bleach to the precipitate, heat to 60-70C stirring periodically, then decant solution and repeat if needed. Red-brown precipitate of hydrated Fe2O3 is produced.

Step 2: prepare individual K and Na trioxalatofertates (my procedure is lengthy, K trioxalatoferrate has relatively low solubility and there are easier preparations)

- Dissolve Fe(OH)3 in oxalic acid, taking 3 mols of acid per 1 mol of Fe(OH)3. Dark brown solution is obtainer.
- Carefully add NaHCO3 or K2CO3 to the solution, again 3 mols of K or Na per 1 mol of Fe. Solution gradually lightens and turns green. Excess would make solution reddish, it can be fixed by additional oxalic acid.
- Evaporate the solution partially, and then cool it down and harvest green crystals of K or Na trioxalatoferrates.

Step 3: mix K and Na salts in 2:1 ... 3:1 molar ratio, prepare saturated solution and grow crystals in any convenient way (K:Na ratio in the salt is 5:1, but pure K salt crystallizes from 5:1 solution, excess of Na is required).